
CHEMISTRY

5070/22

Paper 2 Theory

October/November 2016

MARK SCHEME

Maximum Mark: 75

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge will not enter into discussions about these mark schemes.

Cambridge is publishing the mark schemes for the October/November 2016 series for most Cambridge IGCSE[®], Cambridge International A and AS Level components and some Cambridge O Level components.

Page 2	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
Section A		
A1(a)	SO ₂ /sulfur dioxide (1)	1
A1(b)	C ₃ H ₈ /propane (1)	1
A1(c)	ZnSO ₄ /zinc sulfate (1)	1
A1(d)	CO/carbon monoxide (1)	1
A1(e)	Na ₃ PO ₄ /sodium phosphate (1)	1
	Total:	5

Question	Answer	Marks
A2(a)	NH ₄ ⁺ (1) SO ₄ ²⁻ (1)	2
A2(b)	warm with sodium hydroxide (1) (damp red) litmus turns blue/ammonia produced turns red litmus blue (1)	2
A2(c)	OH ⁻ + H ⁺ → H ₂ O (1)	1
	Total:	5

Page 3	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
A3(a)	<p>One mark each for any two of :</p> <ul style="list-style-type: none"> • same functional group • (same) general formula • similar chemical properties/react similarly • trend in physical properties • each differs by CH₂ 	2
A3(b)	$ \begin{array}{cccc} \text{H} & \text{H} & \text{H} & \text{H} \\ & & & \\ \text{H}-\text{C}-\text{C}-\text{C}-\text{C}-\text{H} & (1) & \text{H}-\text{C}-\text{C}-\text{C}-\text{H} & (1) \\ & & & \\ \text{H} & \text{H} & \text{H} & \text{H} \\ & & & \\ & & \text{H}-\text{C}-\text{H} & \\ & & & \\ & & \text{H} & \end{array} $	2
A3(c)(i)	<p>One mark each for any two of:</p> <ul style="list-style-type: none"> • (generally) increases as number of carbon atoms increases • increases in zigzag way/idea of going up irregularly/idea of going up unevenly • decreases between C2 and C3 compound (or words to that effect) 	2
A3(c)(ii)	any value between –25 (°C) and –45 (°C) (inclusive of these values) (1)	1
A3(d)(i)	<p>C₅H₁₂ + 8O₂ → 5CO₂ + 6H₂O correct formulae (1) correct balance dependent on correct formulae (1)</p>	2

Page 4	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
A3(d)(ii)	Any two of: carbon monoxide/carbon/water (1) carbon monoxide is toxic/carbon monoxide is poisonous (1)	2
	Total:	11

Question	Answer	Marks
A4(a)	propane: increases (rate)/doubling concentration doubles rate (1) iodine: no effect (on rate) (1) hydrochloric acid: increases (rate)/doubling concentration doubles rate (1)	3
A4(b)	particles move faster/particles have more energy (1) more particles have (energy greater than) the activation energy/more successful collisions (1)	2
A4(c)	<u>atom(s)</u> with same <u>number</u> of protons and different number of neutrons/ <u>atom(s)</u> of the same element with different <u>number</u> of neutrons (1)	1
A4(d)	$I_2 + 2At^- \rightarrow At_2 + 2I^-$ (1)	1
A4(e)	they lose electrons/they are oxidised/they give electrons (to hydrogen peroxide)/the oxidation number of iodine increases/the oxidation number of oxygen is decreased (1)	1
	Total:	8

Page 5	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
A5(a)	34.5% (2) If two marks not scored, 171 (for molar mass of nickel carbonyl) scores 1 mark	2
A5(b)	mol nickel carbonyl = $\frac{1.71}{171}$ OR 0.01 (mol) (1) mol gases = (0.01) × 5/idea of multiplying mol × 5 (1) volume of gases = 1.2 dm ³ /1200 cm ³ (units must be correct) (1)	3
A5(c)	low melting point/low boiling point (1) does not conduct electricity (1)	2
A5(d)	<u>isotopes</u> (1) cobalt has greater proportion of heavier isotopes than nickel/nickel has lower proportion of lighter isotopes than nickel (1)	2
A5(e)	acid that is completely ionised/acid that is completely dissociated/acid that releases all ionisable hydrogen (1)	1
	Total:	10

Page 6	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
A6(a)	electrons move (throughout the structure)/mobile electrons (1)	1
A6(b)	One mark each for any two of: <ul style="list-style-type: none"> • decrease in melting point down the group, Or Reverse Argument (ORA) /decrease in boiling point down the group ORA (1) • increase in density down the group ORA • increase in reactivity down the group ORA • decrease in hardness down the group ORA 	2
A6(c)	rubidium hydroxide (1) hydrogen (1)	2
A6(d)	sodium more reactive (than titanium) ORA/sodium higher in reactivity series (than titanium) ORA (1)	1
	Total:	6

Question	Answer	Marks
Section B		
B7(a)	vanadium pentoxide/vanadium(V) oxide/ V_2O_5 (1)	1
B7(b)	four bonding electrons between the two oxygen atoms AND four non-bonding electrons on each oxygen atom (1)	1

Page 7	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
B7(c)(i)	little/no change in % yield at low temperature (1) yield decreases as temperature increases (1) reaction is exothermic so equilibrium moves to the left as temperature increases/increasing temperature favours the endothermic reaction (1)	3
B7(c)(ii)	at 250 °C reaction is slow(er)/at 450 °C reaction is fast(er) (1) idea of compromise temperature/idea of balance between lower yield and faster rate (1)	2
B7(d)	(position of) equilibrium moves to the right/(position of) equilibrium moves towards the product side (1) more (gas) molecules on left than right of the equation/fewer (gas) molecules on right of equation than left (1)	2
B7(e)	$\text{SO}_3 + 2\text{HBr} \rightarrow \text{SO}_2 + \text{Br}_2 + \text{H}_2\text{O}$ (1)	1
	Total:	10

Question	Answer	Marks
B8(a)	reactants labelled on left and products labelled on the right AND product level below reactant level (1) enthalpy change labelled and shown by downward arrow (1) activation energy shown as upward arrow from left hand energy level to energy 'hump' above the highest energy levels of both products and reactants (1)	3
B8(b)	C=C double bond/it is unsaturated (1)	1

Page 8	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
B8(c)	$\begin{array}{c} \text{CN} \\ \\ \text{CH} - \text{CH}_2 \end{array}$ (1) extension bonds shown (1)	2
B8(d)(i)	One mark each for any two of: <ul style="list-style-type: none"> • high melting point/high boiling point • high density • hard • good conductor of electricity/good conductor of heat • malleable • ductile • shiny/lustrous 	2
B8(d)(ii)	$\text{Mo(s)} + 3\text{Cl}_2(\text{g}) \rightarrow \text{MoCl}_6(\text{s})$ correct formulae (1) correct state symbols, dependent on correct formulae (1)	2
	Total:	10

Question	Answer	Marks
B9(a)	magnesium loses (outer shell) electrons more easily than copper/copper cannot give (outer shell) electrons to hydrogen ions but magnesium can ORA (1)	1
B9(b)	the more zinc the greater the strength/the more copper, the lower the strength (1)	1

Page 9	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
B9(c)	layers (of atoms/ions) in copper can slide (when a force is applied) (1) atoms/ions of zinc are different size to those of copper/atoms of zinc disrupt the copper lattice (1) layers (of atoms/ions) in alloy cannot slide (as easily) (1)	3
B9(d)	two rods dipping into a liquid and joined to a power supply (1) electrolyte labelled as 'aqueous copper ions'/named solution of soluble copper salt e.g. aqueous copper sulfate/solution of copper sulfate (1) copper anode/positive electrode AND steel cathode/ negative electrode (1)	3
B9(e)	Cu ₂ O (2) If two marks not scored, mol Cu = $\frac{9.86}{64}$ AND mol O = $\frac{1.23}{16}$ scores 1 mark	2
	Total:	10

Page 10	Mark Scheme	Syllabus	Paper
	Cambridge O Level – October/November 2016	5070	22

Question	Answer	Marks
B10(a)	filter (off excess copper oxide) (1) evaporate (filtrate) to crystallisation point/leave in warm place to crystallise/heat (solution) until saturated (1) dry with filter paper/dry in a drying oven (1)	3
B10(b)	mol sulfuric acid = $2.0 \times \frac{15}{1000}$ OR 0.03 (1) molar mass of hydrated copper sulfate = 250 (1) mass of hydrated copper sulfate = 7.5 g (1)	3
B10(c)	(light) blue precipitate (1) dark blue solution (in excess) (1)	2
B10(d)	anode: oxygen/O ₂ (1) cathode: copper/Cu (1)	2
	Total:	10